

# Raman Microspectroscopy as a Unique Method of the Investigation of Acid Proof Steel Foil Oxidation

A. Weselucha-Birczyńska<sup>1</sup>, J. Dutkiewicz<sup>1</sup>, E. Bielańska<sup>2</sup>, L. Lityńska-Dobrzyńska<sup>3</sup>, M. Kozicki<sup>1</sup>,  
W. Łasocha<sup>1</sup>, A. Rafalska-Łasocha<sup>1</sup>, M. Najbar<sup>1\*</sup>

<sup>1</sup>Faculty of Chemistry, Jagiellonian University, Kraków, Poland

<sup>2</sup>Jerzy Haber Institute of Catalysis and Surface Chemistry, Polish Academy of Sciences,  
Kraków, Poland

<sup>3</sup>Institute of Metallurgy and Materials Science Polish Academy of Sciences, Cracow, Poland  
Email: \*mnajbar@chemia.uj.edu.pl

Received May 21, 2013; revised June 25, 2013; accepted July 23, 2013

Copyright © 2013 A. Weselucha-Birczyńska *et al.* This is an open access article distributed under the Creative Commons Attribution License, which permits unrestricted use, distribution, and reproduction in any medium, provided the original work is properly cited.

## ABSTRACT

In this paper, the results of the investigation of the morphology and phase composition of the oxide layers formed on the surface of the 1H18N9T acid proof steel foil by confocal Raman micro-spectroscopy with optical microscopy, SEM, XRD and TEM-EDS-SADP are presented. The foil oxidation was performed by thermo-programmed heating up to 823, 1023 or 1113 K and next annealing at the final temperatures in the air flow for 48 h, 4 h and 4 h, respectively. The great advantages of the use of the Raman spectroscopy for the phase determining in the oxide layers on the acid proof steel foil are shown. Moreover the possibility of applying the optical microscopy for investigation of the surface morphology of both the initial steel foil and the oxide layers is pointed out.

**Keywords:** 1H18N9T Steel Foil; Oxidation; Raman Microspectroscopy; Morphology; Phases

## 1. Introduction

Monolithic catalysts are commonly used for air pollution abatement. At the beginning, they were produced with the use of ceramic monoliths. Nowadays the metallic monolithic three-way catalysts, for NO<sub>x</sub>, CO, and HC removal from spark engine exhaust, become competitive to the ceramic ones. The oxidation of the ferralloy (Fe, Cr, Al, Y) or kantal (Fe, Cr, Al, Co) steel foils results in the formation of adhesive alumina layers [1].

Recently our interest is focused on oxidation of the metallic monoliths manufactured from the 1H18N9T acid proof steel foil [2]. The spinel phases and Me<sub>2</sub>O<sub>3</sub> type oxides are formed on channel walls of those monoliths during the oxidation. These phases show good activity in the NO decomposition [3] in the carbon particle presence, thus, they could be used as environmental catalysts for nitrogen oxides removal from off gases of fossil fuel combustors.

The morphology, as well as chemical and phase composition of the surface layers of the oxidized monolith channels, strongly depends on the oxidation temperature [2].

The knowledge of the influence of the conditions of the oxidation on phase composition and morphology of the oxide layers is very important for designing the procedure of the monolithic catalyst production [2,3].

In the present paper, the possibilities of the use of the Raman microspectroscopy with optical microscopy for determination of the morphology and the phase composition of the oxide layers are tested.

The optical microscopic images of the foil surface are compared with secondary electron (SE) images obtained in the scanning electron microscope (SEM).

The Raman spectra of the oxide layers are compared with X-ray diffraction (XRD) patterns and with selected area diffraction patterns (SADP) combined with the results of energy dispersive spectrometry (EDS) obtained in analytical transmission electron microscope (ATEM).

## 2. Experimental

### 2.1. 1H18N9T Acid Proof Steel Foil Oxidation

The foil oxidation was accomplished during thermo-programmed (3 K·min<sup>-1</sup>) foil heating up to the temperatures: 823 K, 1023 K or 1123 K in the air flow (90

\*Corresponding author.

$\text{cm}^3 \cdot \text{min}^{-1}$ ) followed by an annealing at those temperatures for 48 h, and 4 h, respectively.

## 2.2. Raman Microspectroscopy with Optical Microscopy

The Oxide phases obtained on the 1H18N9T acid proof steel foils were excited with 785 nm line of the HP NIR diode laser (grating 1200 g/mm) and investigated with Renishaw InVia Raman spectrometer working in confocal mode in a backscattered geometry. A Leica microscope with 50x magnification objective was used to focus a laser on a samples ( $\text{NA} = 0.50$ ). The laser power ( $\sim 1$  mW) was kept enough low to avoid phase modifications within oxide layers. The laser light was focused on the chosen points of the optical image.

## 2.3. Scanning Electron Microscopy

A SEM JEOL JSM-7500F with INCA Oxford Instrument EDX system was used to investigate the morphology of the surface of the initial and oxidized foils.

## 2.4. X-Ray Diffraction

An X'pert PRO MPD diffractometer with 2-d detector PIXCEL and the Bragg-Brentano reflection geometry was used to determine the phases present in the metallic foil and in oxide layers.

## 2.5. Transmission Electron Microscopy

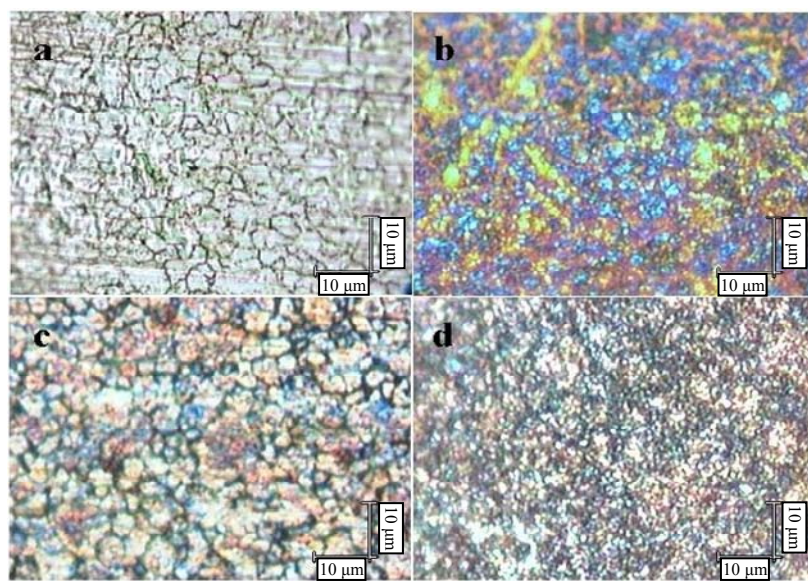
An analytical transmission electron microscope (ATEM) FEI Tecnai G2 at 200 kV equipped with energy dispersive X-ray (EDX) detector was used to take a micro-

scopic images selected area diffraction patterns (SADP) and chemical analysis of the chosen crystallites transparent for electrons. In order to prepare TEM samples, the oxide layer was scraped from foil and then the thin slices of it were placed on a carbon support copper grid.

## 3. Results

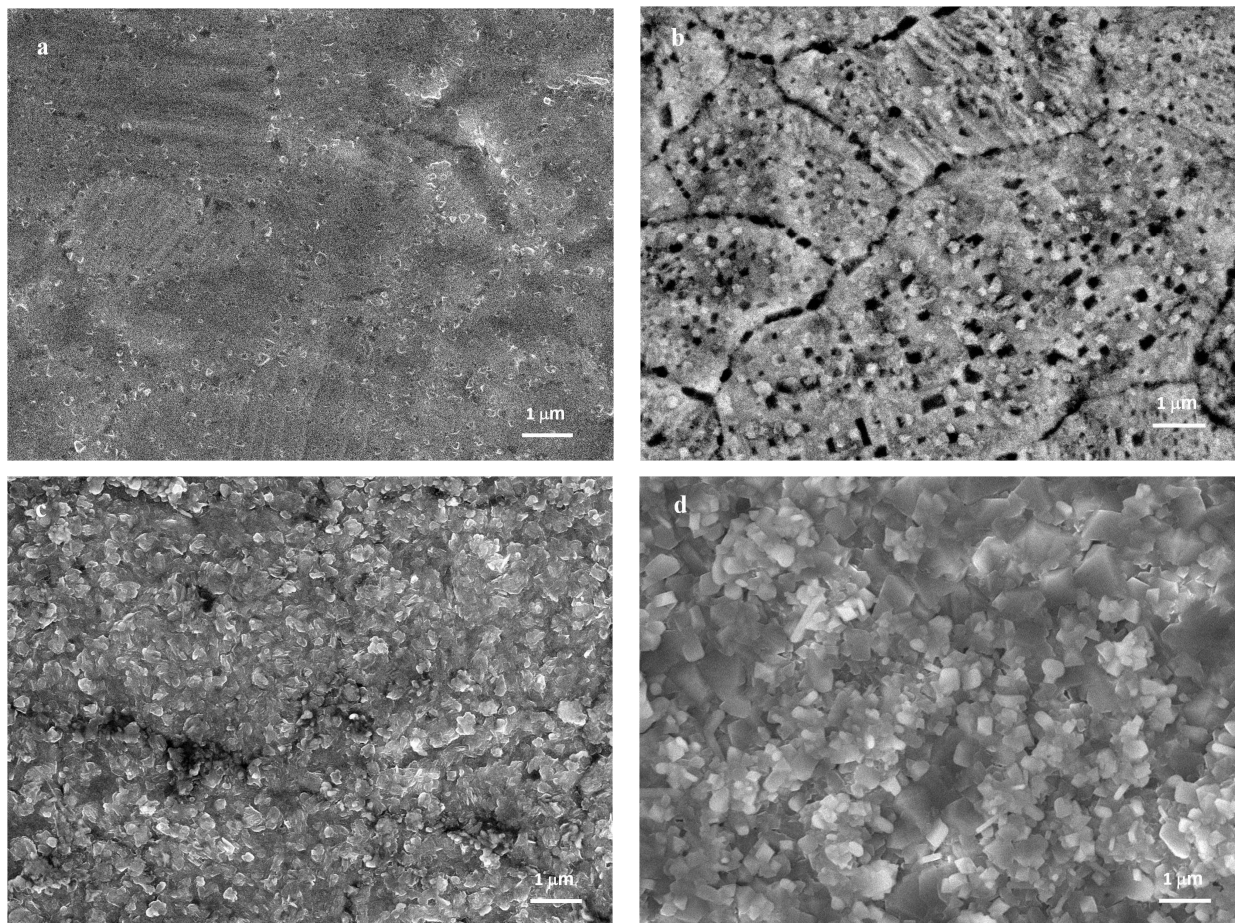
In **Figure 1** the optical images of the initial foil (a) as well as of the foil oxidized up to 823 K (b), 1023 K (c) and 1113 K (d) are presented. In the optical image of the initial foil the metallic grains boundaries are well seen. In the image of the foil oxidised up to 823 K the areas with the colors changing from light brown to blue with the dimensions much greater than those of the metallic grains occur. The dimensions of the oxide areas formed during the oxidation up to 1023 K are similar to the dimensions of the steel grains. Finally, the oxide areas seen in the optical image of the foil oxidised up to 1113 K are of the order of magnitude smaller than those of the foil grains.

The SEM images of the of the initial foil (a) as well as of the foil oxidized up to 823 K (b), 1023 K (c) and 1113 K (d) (magnification 10,000 $\times$ ) are presented in **Figure 2**. In the secondary electron images (SEI) of the initial foil and of the foil oxidised up to 823 K steel grain boundaries are hardly seen. In the SEI of the foil oxidised up to 823 K small particles (possibly crystallites) of the average dimensions of 0.2  $\mu\text{m}$ , inhomogeneously distributed over the surface, are visible. The oxide crystallites formed in the course of the foil oxidation up to 1023 K of average diameters of 0.7  $\mu\text{m}$  strictly cover the surface of the metallic grains. However, the crystallites formed during the oxidation to 1113 K of the dimensions ca 1  $\mu\text{m}$  are



**Figure 1.** Optical images of: initial 1H18N9T steel foil (a); foil oxidized up to 823 K (b); foil oxidized up to 1023 K (c); and foil oxidized up to 1113 K (d).





**Figure 2.** SEM images (10,000 $\times$ ) of: initial 1H18N9T steel foil (a); foil oxidized up to 823 K (b); foil oxidized up to 1023 K (c); and foil oxidized up to 1113 K (d).

separated and spread homogeneously over the whole foil surface. This particular crystallites correspond to the spots in the **Figure 1(d)**.

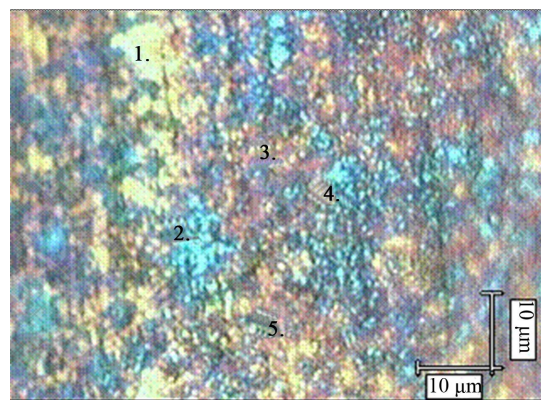
The special attention should be paid to the comparison of the **Figures 1(b)** and **2(b)**. The color differentiation in the **Figure 1(b)** may be caused by inhomogeneous phase distribution or by differences in thickness of the oxide layer in particular areas. The Raman spectral mapping of that foil was performed to elucidate this problem.

In **Figure 3** the points of the confocal spectral analysis were marked on the optical image of the surface foil. The points were chosen on the image areas of different color: 1 and 3—yellow; 2 and 4—blue; and 5—brown.

In **Figure 4**, the spectra from the areas marked by the spots are presented.

Comparison of the spectra taken from the areas marked by 1 - 5 spots allows to find that all of them contain mainly peaks from hematite (295, 415, 228, 500, and 1330  $\text{cm}^{-1}$ ) [4,5] and magnetite (657  $\text{cm}^{-1}$ ) [5,6].

The intensive peak at 616  $\text{cm}^{-1}$  being superposition of the main peak of manganese ferrite at 627  $\text{cm}^{-1}$  [7] and the weak peak of hematite at 610  $\text{cm}^{-1}$  [4,5] shows the



**Figure 3.** Optical image of the surface of the 1H18N9T steel foil oxidized up to 823 K with marked spots of the Raman mapping.

presence of manganese ferrite in all the analysed areas. It can be, thus thought that differentiation of the colors did not arise from the differences in the phase distribution but rather from the differences in the thickness of the oxide layer in particular areas. Such thickness difference may be caused by different steel grain orientations.





**Table 1. XRD pattern of the foil oxidized up to 1113 K and the results of qualitative and semi-quantitative phase analysis.**

Symbol	Ref. Code	Compound Name	Chemical Formula	Space Group	Semi Quant. [%]
*	00-054-0331	410-L Stainless Steel, ferrite	Fe - Cr	1 m - 3 m	74
o	04-008-8265	Tetrateenite	Fe - Ni	P4/mmm	4
v	04-006-8206	Chromium Iron Oxide	Cr <sub>1.4</sub> Fe <sub>0.6</sub> O <sub>3</sub>	R-3c	8
+	04-005-8769	Chromium Iron Oxide	Cr <sub>1.45</sub> Fe <sub>1.55</sub> O <sub>4</sub>	I41/amd	14
	or	or	or	141/amd	
	01-082-0663	Manganese, Chromium Manganese Oxide	Mn(CrMn)O <sub>4</sub>		

phases reflections. Thus, the reliable phase identification is rather impossible.

Since Raman spectroscopy is limited to the oxide phases the problem of the overlapping of the spinel peaks with metal peaks does not exist and the peak from spinel phase in the oxides formed in the course of the oxidation up to 1113 K can precisely be ascribed to the manganese chromite.

In **Figure 7**, the TEM microstructure of the particle of the size about 100 nm, containing mainly Cr, Mn and small amount of Fe is shown (the results of EDX microanalysis were inserted).

Corresponding SADP allow to identify these crystallites as MnCr<sub>2</sub>O<sub>4</sub> spinel. The results of the EDX microanalysis distinctly confirm cationic substitution of Fe for Cr.

However, to prepare the sample for TEM investigation the oxides should be first scratched from the foil, next grounded in the mortar and then placed on the carbon film laying on the copper grid. At last the obtained SADP have to be indexed.

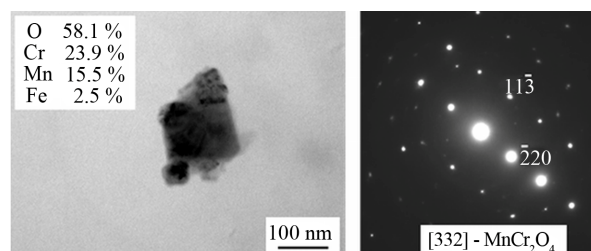
#### 4. Summary

In the presented paper, the possibility of the use of the Raman microspectroscopy with optical microscopy for determination of the morphology and the phase composition of the oxide layers formed during foil oxidation in the air flow is tested.

The comparison of the optical and secondary electron images allows to state that optical microscopy and SEM are complementary methods for determining the morphology of the surface of the oxide layers formed on the acid proof steel foil. The optical images give the reliable information about the dimensions of the steel grains and allow to estimate the influence of their orientations on the oxide layer thickness.

On the other hand, SE images allow to observe crystals growth in the oxide layers.

The comparison of the Raman spectra of the oxide layers with X-ray diffraction patterns and with electron selected area diffraction patterns shows that the confocal



**Figure 7. TEM bright field image of chosen crystallite of the oxide layer formed during oxidation up to 1113 K (EDX result in at.% is inserted) and corresponding SADP.**

Raman microspectroscopy is a unique non-destructive method for the phase identification in the oxide layers on the metallic foil.

#### 5. Acknowledgements

The authors acknowledge the financial support by the European Regional Development Fund under the Innovative Economy Operational Program 2007-2013, POIG. 01.01.02-12-112/09 project.

#### REFERENCES

- [1] J. Camra, E. Bielańska, A. Bernasik, K. Kowalski, M. Zimowska, A. Białas and M. Najbar, "Role of Al Segregation and High Affinity to Oxygen in Formation of Adhesive Alumina Layers on FeCr Alloy Support," *Catalysis Today*, Vol. 105, No. 3-4, 2005, pp. 629-633. doi:10.1016/j.cattod.2005.06.015
- [2] E. Bielańska, J. Camra, J. Dutkiewicz, P. Kornelak, M. Najbar, A. Weselucha-Birczyńska and T. Wilkosz, "Oxidation of Acid-Proof Steel Foil as a Method of Deposition of Phases Active in Nitrogen Oxides Decomposition on Metallic Monolith Walls," *Advanced Materials Research*, Vol. 651, 2013, pp. 317-320. doi:10.4028/www.scientific.net/AMR.651.317
- [3] J. Dutkiewicz, S. Janiga, I. Nazarczuk, M. Kozicki, P. Kornelak and M. Najbar, "Effect of Soot on NO Decomposition in O<sub>2</sub> and SO<sub>2</sub> Presence over Ni-Cr-Fe Oxide Catalysts," *Proceeding of the International Symposium on Nitrogen Oxides Emission Abatement*, Zakopane, 4-7 September 2011, pp. 137-138.
- [4] I. R. Beattie and T. R. Gilson, "The Single-Crystal Ra-

- man Spectra of Nearly Opaque Materials. Iron(III) Oxide and Chromium(III) Oxide,” *Journal of the Chemical Society (A)*, 1970, pp. 980-986.
- [5] D. L. A. de Faria, S. V. Silva and M. T. de Oliveira, “Raman Microspectroscopy of Some Iron Oxides and Oxyhydroxides,” *Journal of Raman Spectroscopy*, Vol. 28, No. 11, 1997, pp. 873-878.  
[doi:10.1002/\(SICI\)1097-4555\(199711\)28:11<873::AID-JRS177>3.0.CO;2-B](https://doi.org/10.1002/(SICI)1097-4555(199711)28:11<873::AID-JRS177>3.0.CO;2-B)
- [6] O. N. Shebanova and P. Lazor, “Raman Spectroscopic Study of Magnetite ( $\text{FeFe}_2\text{O}_4$ ): A New Assignment for the Vibrational Spectrum,” *Journal of Solid State Chemistry*, Vol. 174, No. 2, 2003, pp. 424-430.  
[doi:10.1016/S0022-4596\(03\)00294-9](https://doi.org/10.1016/S0022-4596(03)00294-9)
- [7] R. T. Downs, “The RRUFF Project: An Integrated Study of the Chemistry, Crystallography, Raman and Infrared Spectroscopy of Minerals,” *Proceedings of the 19th General Meeting of the International Mineralogical Association*, Kobe, 23-28 July 2006, pp. 3-13.
- [8] Y. Chen, Z. Liu, S. P. Ringer, Z. Tong, X. Cui and Y. Chen, “Selective Oxidation Synthesis of  $\text{MnCr}_2\text{O}_4$  Spinel Nanowires from Commercial Stainless Steel Foil,” *Crystal Growth & Design*, Vol. 7, No. 11, 2007, pp. 2279-2281. [doi:10.1021/cg070514a](https://doi.org/10.1021/cg070514a)